

Name: _____

CHEM CROSSWORD II

Across

3. The guy who made this puzzle late last night so that you would understand all of this wonderful science.
5. A scale that measures of the concentration of H^+ ions in solution (compared to OH^-) - It is a negative log scale (10^{-x}).
8. Polar molecules clinging to another substance - like water to glass.
10. H^+ ions - nothing but a single, naked proton looking for swinging electrons.
11. OH^- ions - the other part of the water molecule, the base. A very, very reactive ion.
14. Compounds that release OH^- ions into solution - pH greater than 7.0!
16. A mixture in which particles are evenly distributed (in equilibrium).

Down

1. A substance composed of 2 or more elements/compounds that are mixed, but not chemically combined.
2. A mixture of water (or another liquid) and another substance that won't dissolve or reach equilibrium.
4. Compounds that release H^+ ions into water solutions - pH less than 7.0!
6. A weak bond that forms due to the attraction that polar molecules have for one another - like water.
7. pH of 7.0, a pure neutral compound.
9. Polar molecules clinging to themselves - like water molecules.
12. Dissolving/breaking-apart.
13. The substance that is dissolved in a solution.
15. The substance that does the dissolving in a solution - like, well, you know, water.

